



Model studies for the synthesis of the antibiotic lactonamycin and the discovery of new reactions and mechanisms for the construction of substituted heterocycles

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ARTICLE INFO

Article history:

Received 2 March 2010

Received in revised form 6 May 2010

Accepted 13 May 2010

Available online 20 May 2010

Keywords:

Cyclisation

Cascade

Lactonamycin

Antibiotic

New reactions

ABSTRACT

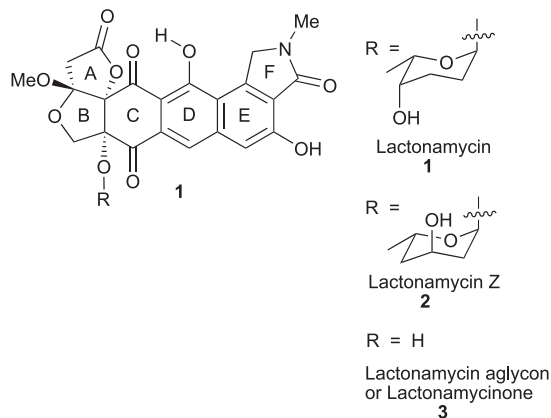
A new and highly efficient route for the construction of a model for the synthesis of lactonamycin **1** is reported. The chemistry has been utilised for the synthesis of heterocyclic rings, and new reactions for the synthesis of dienes and alkynes are reported.

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1. Introduction

Lactonamycin **1** was isolated from bacterium *Streptomyces rishiriensis* and its structure was reported by Matsumoto in 1996.¹ Lactonamycin was shown to possess potent antibiotic activity and good efficacy against Gram-positive bacteria including most notably Methicillin-Resistant *Staphylococcus Aureus* (MRSA) and Vancomycin-Resistant *Enterococcus* (VRE).¹ Matsumoto and co-workers published the structure of lactonamycin **1** in detail in 1999 together with its biological properties and production procedures.^{2,3} In 2003, an analogue of lactonamycin, lactonamycin Z **2** was isolated from pine wood samples collected from Hamsterley Forest in England.⁴ Although lactonamycin Z **2** showed weak Gram-positive activity, it proved to be very active against gastric adenocarcinoma (HMO2) (IC₅₀, 0.19 mg mL⁻¹) in human cell lines.

The biological profile of the lactonamycins is of great interest and because the search for new antibiotics is now of paramount importance due to the dramatic increase in post operative hospital infections⁵ great interest has been shown from the chemical community in the synthesis of the lactonamycins.

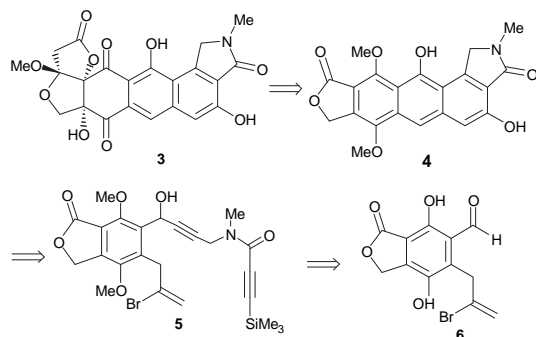


In 2002, Behar and Deville published a synthesis of the ABCD ring system precursor of lactonamycin **1**.⁶ A different approach was published by Kelly and co-workers,⁷ which involved a Diels–Alder reaction as the key step for the formation of the ABCD ring. Kelly and his co-workers subsequently published a route to the EF ring system of lactonamycin **1**⁸ but they have hitherto not reported a total synthesis of lactonamycin **1**. Notable contributions to the synthesis of lactonamycin **1** have also been made by Barrett and Danishefsky.

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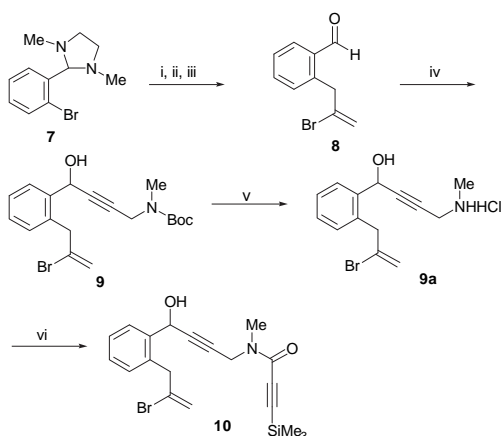
1.1. Cyclisation studies and the new synthetic approach

We devised a novel approach to the core of lactonamycin **1** together with an approach to the total synthesis of this important natural substance⁹ (Scheme 1).



Scheme 1.

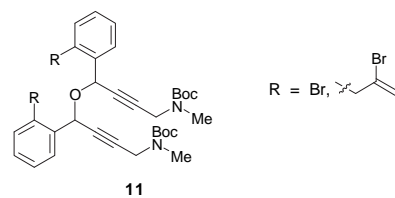
Retrosynthetic analysis of lactonamycin **3** led to the idea that a palladium¹⁰ or tin¹¹ mediated cyclisation of the ene diyne **5** would afford suitable chemical transformations give the tetracycle **4**. We envisaged that the ene diyne **5** could easily be prepared from the aldehyde **6**. In order to test this hypothesis the chemistry outlined in Scheme 2 was carried out.



Scheme 2. Reagents and conditions: (i) *n*-BuLi, (1.05 equiv), THF, $-78\text{ }^{\circ}\text{C}$; (ii) CuCN, (1.05 equiv), $-78\text{ }^{\circ}\text{C}$ to $-40\text{ }^{\circ}\text{C}$; (iii) 2,3-di bromopropene, (1.2 equiv), $-78\text{ }^{\circ}\text{C}$ to room temperature, HCl/H₂O quench then 2 M HCl (73%); (iv) LiC≡C–CH₂NMe(Boc), then ^tBuBr (80%); (v) 4 M HCl dioxane (2 equiv), 2 h (69%); (vi) TMS–C≡C–CO₂H, oxalyl chloride/DMF/Et₃N (86%).

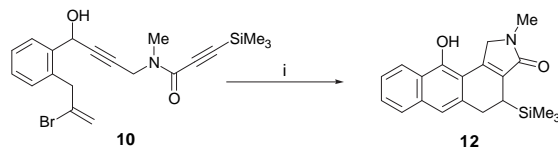
Treatment of the aminal **7** with *n*-butyllithium in tetrahydrofuran at $-78\text{ }^{\circ}\text{C}$, followed by transmetalation with cuprous cyanide¹² gave an intermediate cuprate, which upon reaction with 2,3-dibromopropene, gave the aldehyde **8** after acidic workup. Addition of the lithium salt of *N*-Boc-methylpropargylamine¹³ to the aldehyde **8** followed by workup with *tert*-butyl bromide gave the alcohol **9** in high yield. *tert*-Butyl bromide acts as a bulky proton source and is entirely neutral as a quenching agent.⁹ Other methods that were used to quench this reaction gave lower yields of the desired alcohol **9** and when ammonium chloride in water was used to work up the reaction mixture the ether **11** was isolated.

The alcohol **9** was treated with hydrogen chloride in dioxane to afford the amine hydrochloride salt **9a**. Treatment of the alcohol **9a** with trimethylsilylpropynoyl chloride generated from the parent acid and oxalyl chloride in dimethylformamide gave the desired alcohol **10**.



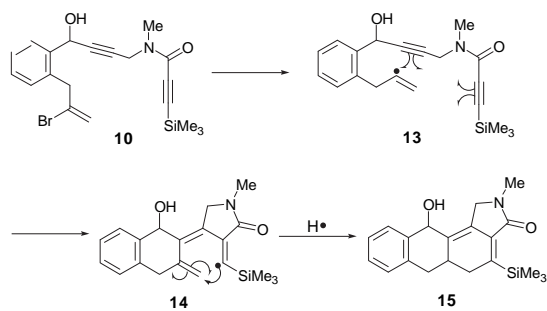
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With the cyclisation precursor **10** in hand, we investigated a radical cascade sequence mediated with tri-*n*-butyltin hydride (Scheme 3).



Scheme 3. Reagents and conditions: (i) Bu₃SnH, AIBN, benzene (14%).

The reaction between tri-*n*-butyltin hydride and the alcohol **10** in boiling benzene gave the tetracyclic amide **12** in 14% isolated yield. This result was surprising because the lactam **12** was not the expected product **15** of the radical cyclisation depicted in Scheme 4.

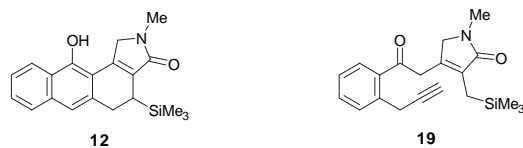


Scheme 4.

When the alcohol **10** was heated in toluene solution in the absence of tri-*n*-butyltin hydride, the tetracyclic lactam **12** was formed in 41% isolated yield. In order to prevent the acid catalysed decomposition of intermediates or the starting material with liberated hydrogen bromide, we elected to carry out the reaction in the presence of the high boiling epoxide epoxyhexene. In the presence of epoxyhexene in hot toluene the alcohol **10** underwent cyclisation to give the amide **12** in 76% yield. By analogy, epoxypropene served as an efficient acid trap in Corey's synthesis of gibberellin acid.¹⁴

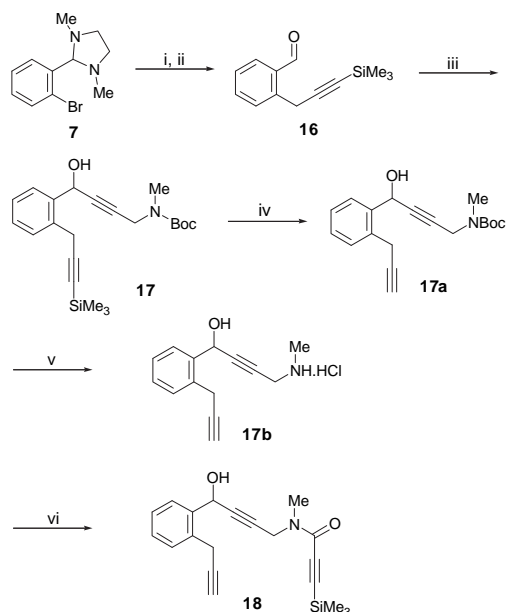
In order to rule out the possibility of an acid catalysed cyclisation the alcohol **18** was prepared according to Scheme 5.

The alkyne **18** was heated in boiling toluene for four hours to give two products **12** and **19**.



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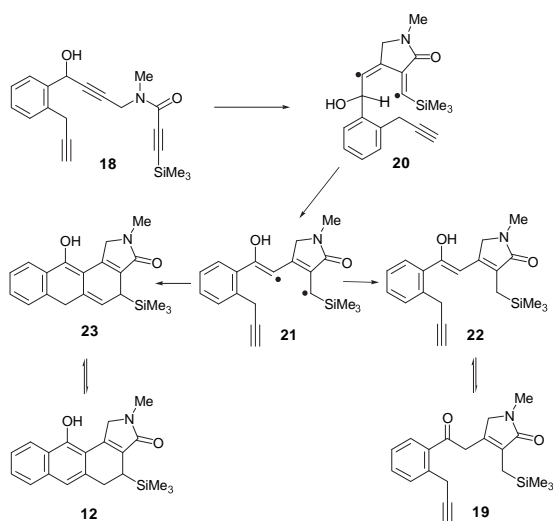
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Scheme 5. Reagents and conditions: (i) *n*-BuLi (1.05 equiv), CuCN (1.05 equiv); Me₃SiC≡C-CH₂Br (1.1 equiv), THF -60 °C (69%); (ii) H₂O; (iii) LiC≡C-CH₂NMe(Boc), ^tBuBr, -95 °C (90%); (iv) MeONa (0.3 equiv), MeOH, DCM 1:1 (97%); (v) HCl, Et₂O (60%); (vi) oxalyl chloride (1.1 equiv), DMF, TMS-C≡C-CO₂H (28%).

1.2. A mechanistic discussion and the discovery of new reactions

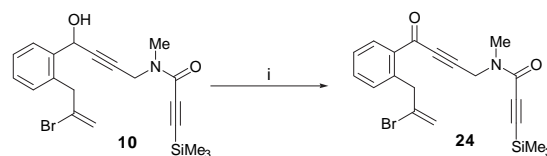
The isolation of the ketone **19** was very interesting because this indicated that a radical mechanism could be in operation. The isolation of **12** in low yield in the presence of tri-*n*-butyltin hydride also points towards a radical cyclisation (Scheme 6).



Scheme 6.

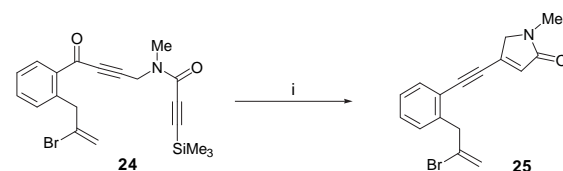
The alkyne **18** could cyclise to give a concentration of the biradical **20**, which in turn could give the biradical **21** by hydrogen atom abstraction. The biradical **21** could then cyclise to give the intermediate **23** and hence isomerise to the phenol **12** or by hydrogen atom abstraction from the solvent (toluene) from the enol **22**, which would form the ketone **19** by tautomerism. We decided to test the mechanism of this intriguing reaction further by

removing the benzylic hydrogen in **10** by benzylic oxidation to afford the ketone **24** (Scheme 7).



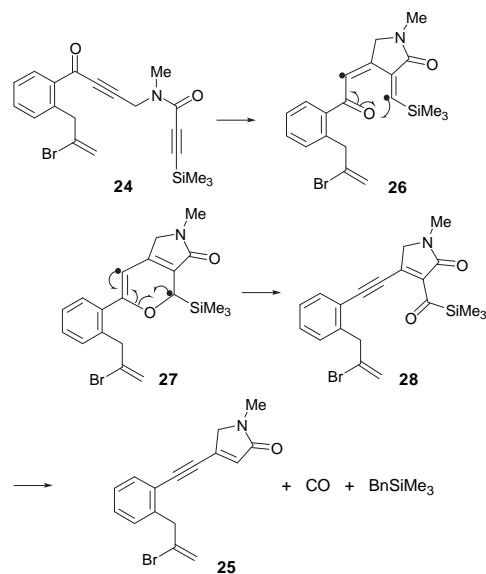
Scheme 7. Reagents and conditions: (i) MnO₂ (21 equiv), DCM (97%).

When the ketone **24** was heated in toluene solution for 6 h under reflux a remarkable transformation was observed; the alkyne **25** was isolated in 30% isolated yield (Scheme 8).



Scheme 8. Reagents and conditions: (i) Toluene, reflux, 6 h (30%).

The formation of the alkyne **25** was totally unexpected and the mechanism outlined in Scheme 9 was proposed to account for this unprecedented transformation.



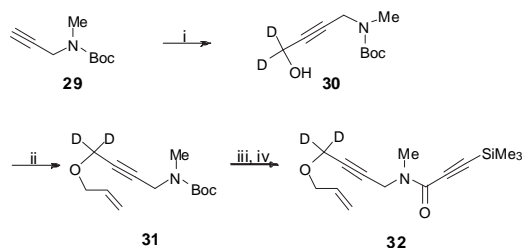
Scheme 9.

The biradical **26** could be formed as described in Scheme 6; instead of hydrogen atom abstraction the alkenyl radical could add to the carbonyl group as shown in structure **26** to form the new biradical **27**. Fragmentation of **27** would give the acyl silane **28**, which on loss of carbon monoxide and benzylsilane (from benzyl radical addition to silicon) would give the observed product **25**.

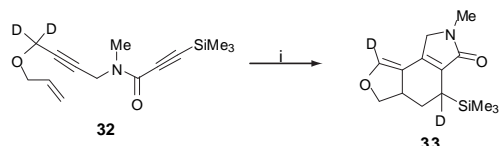
In order to further investigate the cyclisation the deuterated amide **32** was prepared according to Scheme 10.¹⁵

When the amide **32** was heated in hot toluene, cyclisation occurred to give the dihydrofuran **33** in 94% isolated yield (Scheme 11).

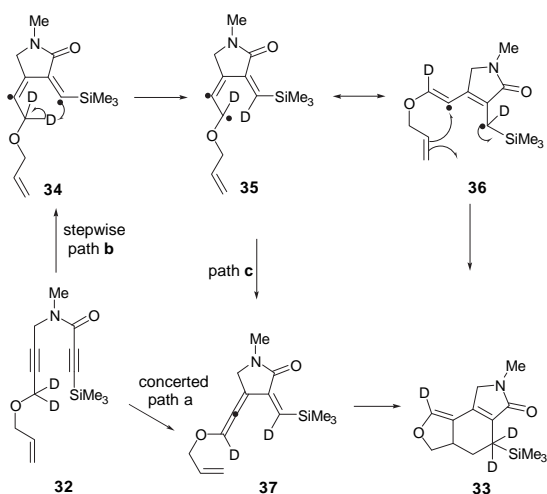
The deuterium transfer observed in lactam **33** coupled with a slower rate of reaction compared with its hydrogen analogue adds support to a concerted mechanism, which could operate in the absence of a radical stabilizing aromatic ring (Scheme 12).



Scheme 10. Reagents and conditions: (i) *n*-BuLi, THF, $-78\text{ }^{\circ}\text{C}$, paraformaldehyde D_2 (94%); (ii) NaH, allyl bromide (quant); (iii) TFA, CH_2Cl_2 ; (iv) Et_3N , $\text{TMSC}\equiv\text{CCOCl}$ (68%).



Scheme 11. Reagents and conditions: (i) Toluene, $110\text{ }^{\circ}\text{C}$, 3.5 h, (94%).



Scheme 12.

In **Scheme 12** concerted path a would proceed by an intramolecular ene reaction to give the allene intermediate **37**, which could also result from the radical pathway b. Either a Diels–Alder reaction from allene **37** to lactam **33** could occur, or the cycloaddition pathway from the biradical **36** would also result in the formation of the lactam **33**.^{16–20}

2. Conclusion

We have discovered a new cyclisation reaction together with a novel route to substituted alkynes. The mechanism of the cyclisation of the aryl substituted diynes appear to follow a radical pathway, which is also postulated for the formation of the alkyne **25**. The facile formation of the lactam **33** occurs with deuterium transfer and the cyclisation requires a longer reaction time (3.5 h) compared to less than one hour for its hydrogen analogue. The formation of **33** is consistent with a radical or a concerted pathway.

3. Experimental

3.1. General

Reactions were conducted at room temperature under an atmosphere of nitrogen unless otherwise stated. They were

monitored using analytical thin layer chromatography with visualisation by UV light and either alkaline potassium permanganate (KMnO_4), vanillin or phosphomolybdic acid (PMA) dips. Reaction solvents were purified and dried according to literature methods. Tetrahydrofuran and diethyl ether were distilled from sodium with benzophenone as indicator; dichloromethane and acetonitrile were distilled from calcium hydride. Petrol refers to distilled petroleum $40\text{--}60\text{ }^{\circ}\text{C}$. All other solvents and reagents were used as supplied. Flash chromatography was performed using silica gel 60, 230–400 mesh. ^1H NMR spectra were recorded on a Bruker 300 MHz machine operating at ambient probe temperature using an internal deuterium lock. Chemical shifts were reported in parts per million (ppm) using residual solvent as an internal standard. Standard abbreviations were used throughout (s singlet; br d broad doublet, br s broad singlet; d doublet; dd doublet of doublets; dt doublet of triplets; dq doublet of quartets; t triplet; q quartet; m multiplet). Coupling constants were measured in hertz (Hz). ^{13}C NMR spectra were recorded at 75 MHz. Chemical shifts were reported in parts per million (ppm). Dept and correlation experiments were used for assignment of spectra. In some cases amide splitting caused doubling of peaks. Suppression of this was not possible using conventional variable temperature experiments and as such the relevant peaks have been assigned together. ESI mass spectra were recorded on a Bruker Daltonics Apex III spectrometer with methanol as solvent. EI mass spectra were recorded on a Fisons VG Autospec spectrometer. Infra red spectra were recorded on a Perkin ELMER Spectrum One FT-IR spectrometer. Crystal structures were obtained from a Bruker/Enraf Nonius FR590 KappaCCD. Structures were named using the ACD labs 'ACD/IUPAC Name v8.05'. The result of ACD/IUPAC Name v8.05 was obtained using the ACD/I-Lab service.

3.1.1. 2-(2-Bromophenyl)-1,3-dimethylimidazolidine (7). A solution of 2-bromobenzaldehyde (10.1 g, 55 mmol) and *N,N'*-dimethylethylenediamine (7.0 ml, 66 mmol) in ethanol (100 ml) was stirred at room temperature for 18 h. The reaction mixture was dried over MgSO_4 and the solvent was removed under reduced pressure to give a yellow oil. This was purified using Kugelrohr distillation (0.2 mbar, $110\text{ }^{\circ}\text{C}$) to yield **7** (12.9 g, 92%) as a transparent colourless oil. Spectra were consistent with those previously reported. IR (neat) ν (cm^{-1}): 3396, 3134, 3064, 2970, 2942, 2840, 2778, 2695, 2671, 2627, 2566, 2473, 1670, 1589, 1568; δ_{H} (300 MHz, CDCl_3): 7.72–7.15 (dd, 1H, $J_1=7.8\text{ Hz}$, $J_2=1.7\text{ Hz}$, Ar), 7.51–7.48 (dd, 1H, $J_1=8.0\text{ Hz}$, $J_2=0.9\text{ Hz}$, Ar), 7.35–7.30 (m, 1H, Ar), 7.16–7.11 (m, 1H, Ar), 4.05 (s, 1H, 7-CH), 3.40–3.34 (m, 2H, 8- CH_2), 2.65–2.60 (m, 2H, 9- CH_2), 2.21 (s, 6H, 10- CH_3 -11- CH_3). δ_{C} (75 MHz, CDCl_3): 138.6 (q, Ar), 132.1 (CH, Ar), 129.6 (CH, Ar), 127.8 (CH, Ar), 125.4 (q, Ar), 88.3 (CH, 7-C), 53.4 (CH_3 , 8-C-9-C), 39.4 (10-C-11-C). MS (EI) m/z (rel int.): 255 (M), 210, 132, 88, 42. HRMS (ESI⁺): calcd for $\text{C}_{11}\text{H}_{15}\text{N}_2\text{Br}$, m/z 255.0491, found 255.0499.

3.1.2. 2-(2-Bromoprop-2-en-1-yl)benzaldehyde (8). *n*-Butyllithium (2.5 M in hexanes, 16.5 ml, 41.2 mmol) was added to a cooled ($-78\text{ }^{\circ}\text{C}$) solution of 2-(2-bromophenyl)-1,3-dimethylimidazolidine (**7**) (10.0 g, 39.2 mmol) in THF (100 ml) and the reaction mixture stirred for 20 min. Copper cyanide (3.7 g, 41.2 mmol) was added and the reaction mixture allowed to warm to between $-45\text{ }^{\circ}\text{C}$ and $38\text{ }^{\circ}\text{C}$ and stirred for 90 min. The reaction mixture was again cooled to $-78\text{ }^{\circ}\text{C}$ and 2,3-dibromopropene (4.9 ml, 47.0 mmol) was added. The reaction mixture was allowed to warm slowly to room temperature, quenched with saturated ammonium chloride solution (100 ml) and extracted with diethyl ether ($2\times 100\text{ ml}$). The combined organic phases were washed with brine, HCl (2 M, $4\times 40\text{ ml}$) with vigorous shaking, saturated sodium hydrogen carbonate solution and brine, dried over magnesium sulfate and the solvent

